



## INTERNATIONAL APPLICATION PUBLISHED UNDER THE PATENT COOPERATION TREATY (PCT)

<b>(51) International Patent Classification <sup>5</sup> :</b> <b>C07C 53/126, 53/128, 51/41</b>	<b>A1</b>	<b>(11) International Publication Number:</b> <b>WO 92/06945</b> <b>(43) International Publication Date:</b> 30 April 1992 (30.04.92)
<b>(21) International Application Number:</b> PCT/EP91/02009 <b>(22) International Filing Date:</b> 18 October 1991 (18.10.91) <b>(30) Priority data:</b> 9022805.7 19 October 1990 (19.10.90) GB <b>(71) Applicants (for all designated States except US):</b> EXXON CHEMICAL LIMITED [GB/GB]; Arundel Towers, Portland Terrace, Southampton, Hampshire SO9 2GW (GB). EXXON CHEMICAL PATENTS INC. [US/US]; 200 Park Avenue, Florham Park, NJ 07932 (US). <b>(72) Inventors; and</b> <b>(75) Inventors/Applicants (for US only) :</b> SARGINSON, Nigel, James [GB/BE]; Constantin Meunierlaan 1, B-3090 Overijse (BE). GODWIN, Allen, David [US/SG]; 335 Bukit Timah Road, #07-03 Wing on Life, Singapore 1025 (SG). REYNIERS, Sylvain, Leontina, Edmond [BE/BE]; Peter Benoitlaan 2, B-1880 Kapelle-op-den-Bos (BE).		<b>(74) Agents:</b> BAWDEN, Peter, Charles et al.; Exxon Chemical Limited, Exxon Chemical Technology Centre, P.O. Box 1, Abingdon, Oxfordshire OX13 6BB (GB). <b>(81) Designated States:</b> AT (European patent), BE (European patent), CA, CH (European patent), DE (European patent), DK (European patent), ES (European patent), FR (European patent), GB (European patent), GR (European patent), IT (European patent), JP, LU (European patent), NL (European patent), SE (European patent), US.  <b>Published</b> <i>With international search report.</i> <i>Before the expiration of the time limit for amending the claims and to be republished in the event of the receipt of amendments.</i>
<b>(54) Title:</b> METAL CARBOXYLATES  <b>(57) Abstract</b>  Basic barium carboxylates of branched chain aliphatic monocarboxylic acids having 6 to 9 carbon atoms are disclosed. The carboxylates may be used, <i>inter alia</i> , as stabilizers for polyvinyl chloride.		

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### Metal Carboxylates

The invention relates to basic barium carboxylates, to a process for the preparation of such carboxylates, and to the use of the carboxylates for, inter alia, stabilizing halogen-containing organic polymers.

Many halogen-containing organic polymers, including, for example, polyvinyl chloride (PVC), must be protected against degradation and discoloration if they are to be exposed to high temperatures or strong light for any length of time. At the elevated temperatures which are necessary for thermoplastic processing, PVC is damaged by dehydrochlorination (release of hydrogen chloride), autoxidation and mechanochemical chain scission. Dehydrochlorination is the main reaction during thermal processing and leads to increasing discoloration of the PVC. A major function of PVC stabilizers is binding and neutralization of the released hydrogen chloride. Other functions include substitution of labile chlorine atoms, additions to polyene sequences, destruction of carbenium salts, and prevention of autoxidation.

A variety of stabilizers have been proposed for halogen-containing polymers, including, for example, the basic alkali and alkaline earth metal salts disclosed in U.S. Specification No. 4 665 117 and the barium- and cadmium-containing organic complexes disclosed in U.S. Specification No. 3 194 823. U.S. Specification

No. 4 665 117 is concerned with a process for improving the colour and stability of basic alkali and alkaline earth metal salts prepared from mixtures containing a phenol. In one embodiment of the process, the reaction mixture contains at least one monocarboxylic acid as well as at least one phenol. The monocarboxylic acid may be an aliphatic or aromatic acid, the aliphatic acids generally having 6 to 30 carbon atoms, with the higher fatty acids containing an average of at least about 12 carbon atoms being preferred. The basic complexes disclosed in U.S. Specification No. 3 194 823 contain cadmium as well as barium. The specification states that the excess metal is believed to be present as colloiddally dispersed metallic carbonate, but that otherwise little is known about the chemical composition of the complexes.

Despite the disclosures of the above-numbered specifications, there remains a need for stabilizers which are easy to prepare, which can readily be dispersed in polymeric materials to be stabilized, and which are highly effective stabilizers for those materials.

The present invention provides a basic barium carboxylate wherein the carboxylate groups consist essentially of groups derivable from at least one branched chain aliphatic monocarboxylic acid having 6 to 9 carbon atoms, preferably 7 to 9 carbon atoms.

The carboxylates of the invention may contain minor amounts of carboxylate groups derived from acids which are not branched chain aliphatic C6 to C9 monocarboxylic acids. Preferably no more than 15 wt %, advantageously not more than 10 wt %, and especially not more than 5 wt %, of the carboxylate groups, based on the total weight of the acids from which the carboxylate groups are derived, are derived from acids other than the specified C6 to C9 acids.

A basic barium carboxylate of the invention may be derived from one, substantially pure, branched chain aliphatic monocarboxylic acid having 6 to 9 carbon atoms. Because commercial processes (e.g., the oxo process carried out on hydrocarbon oligomer feedstocks) for the preparation of such acids normally result in a mixture of products, however, the basic barium carboxylates will more normally be derived from a mixture consisting essentially of two or more branched chain aliphatic monocarboxylic acids having 6 to 9 carbon atoms. Where at least two branched chain aliphatic C6 to C9 monocarboxylic acids are present, a first such acid may contain the same number of carbon atoms as a or the second, but be in a different isomeric form from the second, or a first such acid may contain a different number of carbon atoms from a or the second. A mixture of acids containing two or more isomers having a particular number of carbon atoms and at least one acid

having a different number of carbon atoms may, of course, also be used.

The aliphatic chain, including the branching group(s), in the acids from which the basic barium carboxylates of the invention are derivable may contain, or be substituted by, one or more atoms or groups which are inert under the conditions to which the acids and carboxylates are subjected during processing and use, but, advantageously, the aliphatic chain, including the branching group(s), is made up of carbon and hydrogen atoms only. Furthermore, the aliphatic chain, including the branching group(s), is preferably saturated, as a basic barium carboxylate derived from an unsaturated acid could in some circumstances lead to discoloration of a polymer with which it is blended.

The acids from which the basic barium carboxylates are derivable may contain one or more branching groups, which are normally ethyl or methyl groups, on the main aliphatic chain. The branching group(s) may be present at any position on the main chain, although acids containing at least one branching group on a carbon atom other than the  $\alpha$ -carbon atom (the carbon atom adjacent to the carboxyl group) are preferred. For some purposes it is advantageous to have  $\alpha$ -branching in addition to non- $\alpha$ -branching. Thus, for example, acids which are di- $\alpha$ -branched as well as being non- $\alpha$ -branched may give carboxylates with a reduced tendency to "plate-out" of

polymers with which they are blended, that is, with a reduced tendency to migrate to the surface of the polymers during processing.

Examples of acids from which the basic barium carboxylates of the invention are derivable are the aliphatic C8 and C9 acids sold by Exxon Chemical under the trade name "Cekanoic", and the branched C7 and C9 neo acids sold by Exxon Chemical, the neo acids being di- $\alpha$  branched and also having branching elsewhere in the chain. Cekanoic C8 acid comprises predominantly C8 non- $\alpha$  branched isomers, while Cekanoic C9 acid is predominantly 3,5,5-trimethyl hexanoic acid. Other acids which may be used include the mixture of C9 isomers known as isononanoic acid, and neooctanoic acid, the last-mentioned acid having two substituents on the  $\alpha$ -carbon atom and also being substituted elsewhere on the chain.

Mixtures of one or more of the acids, or types of acids, referred to above may of course be used.

The barium carboxylates of the invention are basic, or overbased, that is to say, the quantity of metal present is greater than that predicted by stoichiometry as being required to react completely with the monocarboxylic acid(s). The extent to which a salt is basic or overbased can be expressed in a number of different ways, but in this specification will be expressed as the basicity, that is, the ratio of metal

equivalents to carboxylic acid equivalents. A neutral salt will have a basicity of 1, while a basic or overbased salt will have a basicity of greater than 1. A high basicity is advantageous as this minimizes the weight of any given carboxylic acid (and thus the weight of barium salt) required for the introduction of a specified weight of barium into, for example, a polymer to be stabilized. The basic carboxylates of the invention have been prepared with basicities of 2 or more, for example, up to about 3.

As will be explained in greater detail below, the preparation of the basic barium carboxylates normally involves, inter alia, the reaction of a barium compound with an acidic gas, that is, a gas which upon reaction with water will produce an acid, the preferred gas being carbon dioxide, the reaction taking place in an organic liquid in the presence of (a) the monocarboxylic acid(s) and a barium compound capable of reacting therewith and/or (b) a barium carboxylate obtainable from the acid(s) and barium compound specified in (a). The precise structure of the basic carboxylates is unknown, but it is thought to be micellar, the micelles, which comprise the barium salt of the acidic gas (for example, barium carbonate) surrounded by barium carboxylate molecules, being in the form of a colloidal solution in the organic liquid. When the carboxylate is used as a stabilizer for a polymer, it is this colloidal solution



that is incorporated into the polymer, and the physical properties of the colloidal solution are thus of considerable importance.

The Applicants have surprisingly found that colloidal solutions of the basic barium carboxylates of the invention have, in a given organic medium and for a given barium content, significantly lower viscosities than colloidal solutions of neutral barium salts of C6 to C9 branched chain monocarboxylic acids and colloidal solutions of barium carboxylates having a greater number of carbon atoms. The viscosities are also significantly lower than those of certain other overbased barium salts previously used as stabilizers for polyvinyl chloride, for example, an overbased barium nonylphenolate containing 27.5 wt % barium. The relatively low viscosities of colloidal solutions of the carboxylates of the invention give rise to a number of advantages, including a saving in production costs because filtration, which is frequently a problem with overbased salts, is easier and quicker, better mixing and processability when the carboxylates are used in stabilizer formulations, including improved filterability if the formulations are to be filtered, and improved dispersion in a polymer, for example, polyvinyl chloride, to be stabilized.

The reduction in viscosities obtainable in accordance with the invention is illustrated in Table I below, where the basicity and viscosity of colloidal

solutions in dearomatized white spirit of a number of different barium salts are given. In the table, the barium content of each colloidal solution is given as a percentage by weight, based on the weight of the colloidal solution.

TABLE I

Acid from which barium salt derived	Barium Content wt%	Basicity	Viscosity at 20°C $10^{-6} \text{ m}^2/\text{s}$
Oekanoic C8	13	1	441
2-Ethylhexanoic	12.5	1	327
Oekanoic C8	16	2	3.5*
Neodecanoic	16	2	17.4
Oekanoic C8	20	2.5	4.9*
Neodecanoic	20	2	36
Oleic	34	>1	332 (at 40°C)

\*carboxylate according to the invention.

By way of further comparison, an overbased barium nonyl phenolate having a barium content of 27.5 wt % had a viscosity of  $2300 \times 10^{-6} \text{ m}^2/\text{s}$  at 25°C.

It will be seen from Table I that the viscosities of the colloidal solutions of the overbased salts are substantially lower, by a factor of about 10 to 100, than those of neutral salts of similar acids, and, for the

overbased salts, the carboxylates according to the invention have lower viscosities, for a given barium content, than the carboxylates of the closely similar C10 branched acid, neodecanoic acid. As discussed elsewhere in this specification, a further advantage of the salts of the invention over the corresponding overbased salts of branched C10 acids is the fact that the salts of the invention having a high barium content can be obtained directly, without the need for a concentration step, and higher basicities are in some cases obtainable with the acids used in accordance with the invention.

As well as the advantages arising from their low viscosities and high barium contents, the basic salts of the invention also have the advantage that they may have low colour, which is particularly significant where white or light-coloured polymers are to be stabilized.

Any suitable process may be used for preparing the basic barium carboxylates of the invention. As indicated above, such a process will normally involve the reaction of a basic barium compound with an acidic gas, especially carbon dioxide, in an organic medium containing (a) the monocarboxylic acid(s) and a barium compound capable of reacting with the acid(s) and/or (b) a barium carboxylate obtainable by the reaction of the acid(s) and barium compound specified in (a), and the invention accordingly also provides such a process. A promoter, preferably a hydroxylic group-containing promoter, for example, a

phenol or, especially, an alcohol, is advantageously also present in the reaction medium.

Processes of the above general type are well known and are described in, for example, the above-mentioned U.S. Specifications Nos. 3 194 823 and 4 665 117, the disclosures of which are incorporated by reference herein. In view of this, it is believed unnecessary to discuss the process of the invention in detail, although process features that have been found to be particularly preferred when preparing the basic barium carboxylates of the invention are discussed briefly below.

Advantageously, the process involves the use of a single barium compound which is capable of reacting with the monocarboxylic acid(s) and with the acidic gas. Suitable barium compounds include barium oxide, barium sulphide, and barium hydrosulphide, but the preferred barium compound is barium hydroxide, which is advantageously used in the form of the octahydrate. Preferably, the reaction medium contains sufficient water to dissolve the barium compound, for example, the barium hydroxide octahydrate.

Preferred promoters for use in preparing the basic barium carboxylates of the invention are polyhydric alcohols, for example, glycols. Triethylene glycol is particularly preferred. A further promoter, for example, a phenol, may also be present, but is not essential. Further, methanol, used as a promoter in many prior

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overbasing processes, is not required, although its use is not excluded. The promoter is advantageously added to a mixture of the barium compound, water, and an organic liquid before introduction of the monocarboxylic acid. The amount of promoter used is not critical, but is advantageously in the range of from 3 to 15 wt%, based on the final colloidal solution.

Any organic liquid which is inert under the conditions used in the process may be used as the medium in which to carry out the reaction. As the overbased carboxylates are used without removal of the organic liquid, the choice of liquid will be dependent in part on the intended end use of the product. Preferably, a hydrocarbon liquid is used, for example, dearomatized white spirit, although an oxygen-containing liquid compound may be appropriate in some circumstances.

The branched aliphatic C6 to C9 monocarboxylic acid(s) used in accordance with the invention is/are advantageously introduced gradually into a mixture comprising the basic barium compound, the promoter, the organic liquid, and optionally water, whereupon the barium compound and acid(s) are thought to react to form the corresponding barium carboxylate(s). (If desired, barium carboxylate(s) could be introduced as such into the reaction medium, rather than being prepared in the medium, but it is normally more convenient to prepare the carboxylates in situ.) It is preferred that the mixture to which the acid is added be at elevated temperature, preferably in the range of 80 to 120°C.

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The reaction medium into which the acidic gas is passed contains an amount of basic barium compound in excess of that required to react with the monocarboxylic acid(s). The acidic gas may be, for example, sulphur dioxide, sulphur trioxide, carbon disulphide or hydrogen sulphide, but is advantageously carbon dioxide. For convenience, the invention will hereafter be discussed with reference to carbon dioxide, but it will be appreciated that, where appropriate, this may be replaced by another acidic gas. Additional barium compound may, if desired, be added during carbonation, but preferably all the barium compound is introduced into the reaction vessel at the outset.

Advantageously, water present in the reaction medium is substantially completely eliminated, for example, by heating, before carbonation is commenced. Carbon dioxide is preferably passed through the reaction medium until no more is absorbed by the reaction medium, carbonation advantageously being effected at an elevated temperature, preferably a temperature of from 130 to 170°C, more preferably 140 to 170°C. The product obtained may then be filtered to remove any unreacted solids.

One or more steps to reduce any tendency to discoloration of the basic salts are advantageously included in the process. Thus, the process is preferably carried out, at least when carbonation is not occurring, under an atmosphere of an inert gas, for

example, nitrogen. Further, an agent to improve the stability of the colloid, for example, an agent that will react with any free hydroxyl groups, is advantageously added to the final product. Preferred agents for stabilizing the colloid are alkyl phosphates, for example, triethyl phosphate.

Normally barium will be the only metal present in the carboxylates of the invention. For some purposes it may be desirable that zinc be present in addition to barium in, for example, a polymer to be stabilized, and, where this is the case, a barium carboxylate may be mixed with a zinc salt before incorporation into the polymer, or a zinc salt may be blended into the polymer separately. In some circumstances, however, it may be advantageous to incorporate zinc into the barium salts of the invention, for example, by replacing some of the barium compound(s) used as starting material by a zinc compound that will react with the carboxylic acid(s) in the reaction mixture used for preparing the overbased barium carboxylates, and the invention accordingly also provides a basic barium carboxylate according to the invention in the form of a complex which also contains zinc. For some purposes another metal, other than cadmium, may replace the zinc in the mixtures and complexes referred to above.

In accordance with the invention, it is possible, especially when the preferred process features referred

to above are used, to obtain colloidal solutions of basic barium salts of C6 to C9 branched chain aliphatic monocarboxylic acids with a relatively high barium content (for example, 20 wt% or more) directly, without the use of a final concentration step to reduce the amount of organic liquid present. As indicated earlier, if the acid(s) used in accordance with the invention are replaced in whole or in substantial amounts by a higher acid, for example, a branched chain C10 aliphatic monocarboxylic acid, such barium contents can only be obtained by a process of the type described if a final concentration step is used.

As indicated above, the basic barium carboxylates of the invention may be used for, inter alia, reducing degradation and/or discoloration of halogen-containing polymers, for example, vinyl chloride polymers, which are exposed to high temperatures and/or light, for any length of time. Accordingly, the invention also provides a polymer composition which comprises a halogen-containing organic polymer, especially a vinyl halide polymer, and a basic barium salt in accordance with the invention. The invention also provides the use of a basic barium carboxylate in accordance with the invention for stabilizing a halogen-containing organic polymer, especially a vinyl halide polymer. The barium carboxylates of the invention are especially useful for stabilizing homopolymers and copolymers of vinyl



chloride, particularly flexible vinyl chloride polymers.

In general, the basic salts are added to the polymer in such an amount that the proportion of barium in the composition is in the range of from about 0.01 to 0.5 wt % barium, preferably 0.025 to 0.25 wt % barium, based on the weight of the polymer and any other components, including plasticizer, that may be present.

Because the use of the specified C6 to C9 acids in the production of overbased barium salts makes it possible to produce salts having a higher proportion of barium than do the corresponding neutral salts, the same barium content can be obtained in a given weight of polymer composition using a lower weight of barium-containing additive than that necessary when using neutral salts. This makes it possible, if desired, to include other components in a stabilizer formulation to be added to a polymer. The invention therefore also provides a stabilizer formulation comprising a basic barium carboxylate according to the invention and a further component, for example, a co-stabilizer. Suitable co-stabilizers include phosphites, for example, diphenyl isooctyl phosphite. Such a formulation may contain, for example, 2 to 15 wt % of barium.

The basic barium carboxylates of the invention may also find application as paint dryers, corrosion inhibitors, and as detergent/dispersant additives for lubricating oils. Where desirable for any particular

application, they may, as indicated above, be used in conjunction with other overbased metal salts. Thus, for example, zinc salts may be desirable constituents of stabilizers for PVC.

Examples 1 to 4 illustrate the invention. Examples A and B are provided by way of comparison.

#### Comparative Example A

Preparation of a basic barium salt of neodecanoic acid.

A reactor was charged with 398 g of barium hydroxide octahydrate and 500 g of water and the mixture was heated slowly to 80°C with thorough stirring. 80 g of triethylene glycol and 327 g of dearomatized white spirit (two-thirds of the total quantity used) were then added, and the resulting mixture was heated slowly to 90 to 95°C. A mixture of 215 g of neodecanoic acid and 164 g of dearomatized white spirit were then slowly added over a period of 20 to 30 minutes, following which the mixture was slowly heated to approximately 105°C and maintained at that temperature for about 1 1/4 hours. Thereafter, the mixture was heated slowly to approximately 155°C in order to eliminate any remaining water.

After allowing the mixture to cool to about 140°C, it was carbonated for 25 minutes at a carbon dioxide flow rate of 48 litres/hour. The carbonated mixture was then

allowed to cool to 120°C and filtered using a CUNO 308 Zeta Plus filter. The filtered product had a barium content of about 17 wt%, a basicity of 2, and a viscosity at 20°C of approximately  $15 \times 10^{-6} \text{ m}^2/\text{s}$ . This product was then concentrated by distillation under vacuum to give a product with a barium content of 20 wt % and a viscosity at 20°C of approximately  $36 \times 10^{-6} \text{ m}^2/\text{s}$ . 2 wt % of triethylphosphate, based on the weight of the concentrated product, was then added to improve the stability of the colloidal solution.

All the steps indicated above, apart from carbonation, were carried out under an atmosphere of nitrogen.

#### Comparative Example B

Preparation of a basic barium salt of a mixture of neodecanoic acid and neononanoic acid.

The procedure indicated in Comparative Example A was repeated, replacing the neodecanoic acid by 217 g of a mixture of neodecanoic acid and neononanoic acid (70 parts by weight of C10 acid to 30 parts by weight of C9 acid). A product containing 20 wt % barium produced by concentrating the filtered product by vacuum distillation had a basicity of 2.

**Examples 1 to 4**

Preparation of basic barium carboxylates according to the invention.

The procedure indicated in Comparative Example A was repeated, with the omission of the step of concentrating the filtered product, replacing the neodecanoic acid by the acid or acid mixture indicated in Table II below. In each case a product with a basicity of 2.5 and a barium content of 20 wt % was obtained directly, without the need for the concentration step used in Comparative Examples A and B. In the Table, amounts are given in grams.

**TABLE II**

	Ex.1	Ex.2	Ex.3	Ex.4
Ba(OH) <sub>2</sub> .8H <sub>2</sub> O	459	459	459	459
Decanoic C8 acid	168	-	-	-
Decanoic C7 acid	-	151	-	-
Decanoic C9 acid	-	-	184	-
Neooctanoic acid	-	-	-	168
CO <sub>2</sub>	38	38	38	38
Deaeromatized white spirit	543	561	527	543
Triethylene glycol	50	50	50	50
Water	500	500	500	500

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Example 5

The effect of the product of Example 1 as a stabiliser for polyvinyl chloride was compared with the commercially available overbased barium nonyl phenolate at the same barium level in the formulation. The stabiliser formulation, the polyvinyl chloride formulation and the heat stability data are given in the following TABLE III.

TABLE III

STABILISER COMPOSITION		CONTENT wt%	
Component		<u>Formulation 1</u>	<u>Formulation 2</u>
Solution of mixed barium and zinc salts of oleic and p-tert butyl benzoic acids		57	57
Overbased barium isooctanoate (Cekanoic C8 acid) (20 wt% barium)		-	20
Overbased barium nonyl phenolate (28 wt% barium)		14.3	-
Diphenyl isooctylphosphite		20	20
Diketone costabilizer		3	3
White Spirit		5.7	-
<u>Polyvinyl Chloride Formulation</u>			
Solvic 246GA	100 parts		
Diocetyl phthalate	25 parts		
Expoxidised soya bean oil	1.5 parts		
Stabilizer	2.5 parts		
<u>Heat Stability Performance</u>			
<u>Oven Test at 190°C</u>			
A. Initial colour		+++	+++
B. Colour hold (40-50 mins)		++	++
C. Complete loss of colour		90 mins	90 mins
<u>Dynamic Test (two roll mill) at 180°C</u>			
D. Initial colour		+++	+++
E. Colour hold (20-30 mins)		++	+++
F. Colour hold (40-45 mins)		+	++

+++ = Excellent, ++ = Good, + = Poor

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The test results show that use of the basic barium salts of acids of this invention give significantly better dynamic colour control in comparison with currently used barium salts.

The Oven Test is performed by cutting a roll milled sheet of the polyvinyl chloride into small pieces which are placed in the oven at the indicated temperature on glass plates and after the indicated time are compared with a standard reference.

In the Organic Test the made up compound is processed on a two roll whose rolls are at the indicated temperature for the time indicated and the colour compared with a standard reference.

SUBSTITUTE SHEET

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CLAIMS:

1. A basic barium carboxylate wherein the carboxylate groups consist essentially of groups derivable from at least one branched chain aliphatic monocarboxylic acid having 6 to 9 carbon atoms.
2. A barium carboxylate as claimed in claim 1, wherein at least one acid from which the carboxylate groups are derivable the or each branching group is a methyl or ethyl group.
3. A barium carboxylate as claimed in claim 1 or 2, wherein at least one acid from which the carboxylate groups are derivable there is at least one branching group on a carbon atom other than the  $\alpha$ -carbon atom.
4. A barium carboxylate as claimed in claim 3, wherein the said one acid also has at least one methyl group on the  $\alpha$ -carbon atom.
5. A barium carboxylate as claimed in claim 1, wherein the carboxylate groups consist essentially of groups derivable from a mixture of acids each having 8 carbon atoms and being branched predominantly in non- $\alpha$  positions.
6. A barium carboxylate as claimed in any one of claims 1 to 5, which has a basicity of at least 2.
7. A barium carboxylate as claimed in any one of claims 1 to 6, in the form of a colloidal solution in an inert organic liquid, the colloidal solution having a barium content of at least 20 wt%.
8. A barium carboxylate as claimed in any one of claims 1

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to 7, in the form of a complex which also contains zinc.

9. A process for preparing a basic barium carboxylate as claimed in claim 1, which comprises reacting a basic barium compound with an acidic gas in an organic liquid in the presence of (a) at least one monocarboxylic acid as specified in claim 1 and a barium compound capable of reacting with the said acid and/or (b) a barium carboxylate obtainable by reaction of the acid(s) and barium compound specified in (a).

10. A process as claimed in claim 9, wherein the same barium compound is used for reaction with the monocarboxylic acid(s) as is used for reaction with the acidic gas.

11. A process as claimed in claim 9 or claim 10, wherein reaction of the barium compound with the carboxylic acid(s) is carried out in the presence of sufficient water to dissolve the barium compound.

12. A process as claimed in any one of claims 9 to 11, wherein substantially all the water present in the reaction medium is removed before introduction of the acidic gas.

13. A process as claimed in any one of claims 9 to 20, wherein the acidic gas is carbon dioxide.

14. A process as claimed in any one of claims 9 to 13, wherein the reaction mixture also comprises a zinc compound capable of reacting with the monocarboxylic acid(s).

15. A polymer composition comprising a halogen-containing polymer and a barium carboxylate as claimed in any one of claims 1 to 8.



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16. The use of a basic barium carboxylate as claimed in any one of claims 1 to 8 for stabilizing a halogen-containing polymer.

17. A composition or use as claimed in claim 15 or claim 16, wherein the polymer is a homo- or copolymer of vinyl chloride.

18. A composition or use as claimed in any one of claims 15 to 17, wherein the proportion of barium is from 0.025 to 0.25 wt%, based on the weight of the polymer and any plasticizer that may be present.

# INTERNATIONAL SEARCH REPORT

International Application No

PCT/EP 91/02009

<b>I. CLASSIFICATION OF SUBJECT MATTER</b> (If several classification symbols apply, indicate all) <sup>6</sup>		
According to International Patent Classification (IPC) or to both National Classification and IPC Int.C1.5                      C 07 C 53/126                      C 07 C 53/128                      C 07 C 51/41		
<b>II. FIELDS SEARCHED</b>		
Minimum Documentation Searched <sup>7</sup>		
Classification System	Classification Symbols	
Int.C1.5	C 07 C 53/00                      C 07 C 51/00	
Documentation Searched other than Minimum Documentation to the Extent that such Documents are Included in the Fields Searched <sup>8</sup>		
<b>III. DOCUMENTS CONSIDERED TO BE RELEVANT<sup>9</sup></b>		
Category <sup>o</sup>	Citation of Document, <sup>11</sup> with indication, where appropriate, of the relevant passages <sup>12</sup>	Relevant to Claim No. <sup>13</sup>
X	GB,A, 891858 (S.A. GHENT) 21 March 1962, see page 1, lines 12-23; page 2, lines 67-80; claims 1,4-6 ---	1-5,16
X	GB,A,1005957 (THE CARBORUNDUM CO.) 29 September 1965, see page 2, lines 30-41; page 3, lines 8-29; claim 11 ---	1-5
Y	EP,A,0279493 (SHELL) 24 August 1988, see page 2, lines 40-53; page 3, lines 8-14; page 3, line 58 - page 4, line 3; claims 1,3,6,7,11-14 ---	1-6,9, 13
Y	EP,A,0010807 (AKZO) 14 May 1980, see page 9, line 30 - page 10, line 27; page 12, lines 4,5; page 19, lines 8-14; claims 1-4,24 --- -/-	1-6,9, 13
<div style="display: flex; justify-content: space-between;"> <div style="width: 45%;"> <p><sup>o</sup> Special categories of cited documents : <sup>10</sup></p> <p>"A" document defining the general state of the art which is not considered to be of particular relevance</p> <p>"E" earlier document but published on or after the international filing date</p> <p>"L" document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)</p> <p>"O" document referring to an oral disclosure, use, exhibition or other means</p> <p>"P" document published prior to the international filing date but later than the priority date claimed</p> </div> <div style="width: 45%;"> <p>"T" later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention</p> <p>"X" document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step</p> <p>"Y" document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art.</p> <p>"&amp;" document member of the same patent family</p> </div> </div>		
<b>IV. CERTIFICATION</b>		
Date of the Actual Completion of the International Search		Date of Mailing of this International Search Report
26-02-1992		27. 03. 92
International Searching Authority		Signature of Authorized Officer
EUROPEAN PATENT OFFICE		Maria Peis <i>Maria Peis</i>

III. DOCUMENTS CONSIDERED TO BE RELEVANT (CONTINUED FROM THE SECOND SHEET)		
Category °	Citation of Document, with indication, where appropriate, of the relevant passages	Relevant to Claim No.
A	FR,A,1540384 (UGINE KUHLMANN) 27 September 1968, see page 1, column 2, lines 4-20; page 1, column 2, line 22 - page 2, column 1, line 5; page 2, column 2, lines 25,26; page 3, column 2, lines 50-53; claims 1,2 -----	1-5

**ANNEX TO THE INTERNATIONAL SEARCH REPORT  
ON INTERNATIONAL PATENT APPLICATION NO.**

EP 9102009

SA 52378

This annex lists the patent family members relating to the patent documents cited in the above-mentioned international search report. The members are as contained in the European Patent Office EDP file on 17/03/92  
The European Patent Office is in no way liable for these particulars which are merely given for the purpose of information.

Patent document cited in search report	Publication date	Patent family member(s)	Publication date
GB-A- 891858		LU-A- 37399	
GB-A- 1005957		None	
EP-A- 0279493	24-08-88	JP-A- 63203645	23-08-88
		US-A- 4810398	07-03-89
EP-A- 0010807	14-05-80	CA-A- 1204755	20-05-86
		JP-A- 55081834	20-06-80
		JP-B- 63023980	18-05-88
FR-A- 1540384		BE-A- 714330	16-09-68
		DE-A,B,C 1768298	12-08-71
		FR-E- 94407	14-08-69
		GB-A- 1198556	15-07-70
		LU-A- 55984	15-07-68
		NL-C- 137462	
		NL-A- 6805806	28-10-68
		NL-A- 7205818	25-07-72
		US-A- 3554909	12-01-71